

FORMATION OF γ' -Ni₃Al VIA THE PERITECTOID REACTION: $\gamma + \beta (+ \text{Al}_2\text{O}_3) = \gamma' (+ \text{Al}_2\text{O}_3)$

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nickel aluminides: solidification: partial thermodynamic properties: Knudsen effusion mass spectrometry

Abstract

The activities of Al and Ni were measured using *multi-cell* Knudsen effusion-cell mass spectrometry (multi-cell KEMS), over the composition range 8 – 32 at.%Al and temperature range $T = 1400 - 1750$ K in the Ni-Al-O system. These measurements establish that equilibrium solidification of γ' -Ni₃Al-containing alloys occurs by the eutectic reaction, $L (+ \text{Al}_2\text{O}_3) = \gamma + \beta (+ \text{Al}_2\text{O}_3)$, at 1640 ± 1 K and a liquid composition of 24.8 ± 0.2 at.%Al (at an unknown oxygen content). The $\{\gamma + \beta + \text{Al}_2\text{O}_3\}$ phase field is stable over the temperature range 1633 – 1640 K, and γ' -Ni₃Al forms via the peritectoid, $\gamma + \beta (+ \text{Al}_2\text{O}_3) = \gamma' (+ \text{Al}_2\text{O}_3)$, at 1633 ± 1 K. This behavior is inconsistent with the current Ni-Al phase diagram and a new diagram is proposed. This new Ni-Al phase diagram explains a number of unusual steady-state solidification structures reported previously and provides a much simpler reaction scheme in the vicinity of the γ' -Ni₃Al phase field.

Introduction

The Ni-Al system is the foundation for Ni-base superalloys used in gas-turbine applications. The precipitation of γ' -Ni₃Al (ordered FCC, L1₂) from γ -(Ni) (disordered FCC) gives these alloys their remarkable high temperature mechanical properties. To fully understand these alloys the equilibrium phase boundaries and solution behavior of γ -(Ni), γ' -Ni₃Al, β -NiAl (ordered cubic, B2) and liquid, L , phases need to be accurately determined in the composition range 10 – 35 at.% Al. This need is widely recognized as shown by the large number of detailed experimental studies^[1-18], assessed phase diagrams^[19-21] and various thermodynamic models^[22-25] published over the last 70 years. The solidification behavior of γ' -Ni₃Al and the Ni-Al phase diagram has been modified a number of times over this period, but the consensus has been γ' -Ni₃Al remains stable up to the liquidus and is involved in both eutectic and peritectic transformations. To date much of the discussion has focused on which side of γ' -Ni₃Al and at what temperature the invariant reactions exist.

It is important to note the transformation from γ -(Ni) and β -NiAl to γ' -Ni₃Al is fast^[4-8, 12-15, 17] and the γ' -Ni₃Al phase field broadens significantly and growth of new γ' -Ni₃Al consumes the eutectic and peritectic microstructures on cooling^[4, 8, 19-21]. As a result it is difficult to quench-in the high-temperature microstructures of γ' -Ni₃Al containing alloys and observe them at room temperature. Experimental techniques that provide information about the phase equilibria at temperature are therefore crucial to studying the Ni-Al system. Many studies of this system have used directional solidification (DS), differential thermal analysis (DTA) and differential scanning calorimetry (DSC) techniques to study the high temperature phase equilibria, but it is always hard to determine equilibrium transformation temperatures or phase boundaries under non-isothermal or irreversible conditions. DS involves significant thermal gradient and heat flux across the solidification

front, while DTA and DSC studies involve heating and cooling rates of 2.0 ~ 10 K/min. DS provides important information about steady-state solidification in these alloys, but a number of studies report a range of microstructures that are unexplainable with the current Ni-Al phase diagram (*e.g.*, $\beta + \gamma$ eutectics and unusual γ' / γ' interfaces)^[12-16]. Reducing the ramping rate of DTA and DSC give closer to equilibrium transformations but reduced signal intensity imposes a low limit on the ramp rate^[30]. Of more concern is the disagreement between transformation temperatures observed during heating and cooling (up to 20 ~ 100 K), which was the major reason sighted by Hilpert for including KEMS measurements in their studies^[8, 9]. The KEMS technique is well suited to study equilibrium solidification behavior because isothermal conditions are obtained (T is routinely held within ± 0.5 K over 1–4 hrs) and equilibrium is closely approached. Equilibrium transformations are determined by extrapolating measured equilibrium behavior on either side of a phase boundary or transformation^[27, 28].

Another aspect that is not always adequately considered is the thermodynamic state of the system present in these experiments. To correctly interpret thermodynamic property measurements the equilibrium needs to be described accurately and any unmeasured variables must be identified and all assumptions clearly stated. The typical approach has been to only consider the Ni-Al alloy and ignore the container material (typically Al_2O_3). In this study the boundary of the system is identified as the inner surface of the effusion-cell and Al_2O_3 is included together with the alloy sample, thus the $\{\text{alloy} + \text{Al}_2\text{O}_3 + \text{vapor}\}$ equilibrium describes the thermodynamic state and the Ni-Al-O ternary system is considered^[29, 31, 32]. In these equilibria the alloy phases γ -(Ni), γ' -Ni₃Al, β -NiAl and L are saturated with O, and the Al_2O_3 is saturated with Al and Ni from the alloy. The amount of O saturation in the alloy and Ni and Al in Al_2O_3 are unknown. This removes the implicit assumptions that dissolved O does not influence the solution behavior of the alloy phases and measurements made in the Ni-Al-O system can be applied directly to the Ni-Al system. A review of the phase equilibrium observed in the Al-O^[33] and Ni-O^[34] systems clearly shows the merit of including O and Al_2O_3 in the description of the results and considering them in terms of the ternary system. This paper reports a series of partial thermodynamic property measurements - activities of $a(\text{Al})$ and $a(\text{Ni})$ - made by the vapor pressure method with a multiple Knudsen effusion cell mass spectrometer (*multi-cell KEMS*) in Ni-Al alloys in equilibrium with Al_2O_3 . Particular attention is given to compositions on either side of γ' -Ni₃Al, nominally Ni-23at.%Al and Ni-27at.%Al in the $\{\gamma' + \gamma + \text{Al}_2\text{O}_3\}$ and $\{\gamma' + \beta + \text{Al}_2\text{O}_3\}$ condensed three-phase fields, respectively. These measurements raise some fundamental questions about the accepted solidification behavior of γ' -Ni₃Al, which are considered in detail.

Experimental

Alloy and Sample Preparation

The alloys considered in this study are listed in Table 1. Their compositions were determined by inductively coupled plasma spectroscopy (using a Varian Vista-Pro ICP-OES) and inert gas fusion for nitrogen / oxygen (using a LECO TC-436). All alloys were prepared by repeated arc-melting (5 times) the correct proportions of the pure elements (at least 99.995 wt.% pure) and cast in water-cooled copper molds. After casting all alloys were homogenized in a flowing argon gas atmosphere at 1473K for 6 hours and at 1423K for an additional 24 hours, water quenched and cut into 1 to 2mm thick slices. Directly prior to loading into the effusion cell (and pumping to 10^{-4} Pa) the surface of each alloy slice was removed by grinding with 600-grit SiC paper, roughly cut into 2mm cubes with a metal shear and ultrasonically cleaned in acetone then ethanol. Typically, 0.7 ~ 1.5 g of alloy sample was loaded into each effusion cell. The effusion cells were fabricated from high purity polycrystalline Al_2O_3 (at least 99.5 wt% pure). Prior to these experiments the effusion cells were used numerous times with similar Ni-Al alloys and were cleaned between each use by soaking in hot aqua regia (HNO_3 : HCl , 1:3) and baking in the temperature range, $T = 1800 - 1850\text{K}$ under vacuum ($\sim 10^{-3}$ Pa) for 10 – 12 hours.

Table 1. Measured Alloy Composition (at.%)

Alloy	Ni	Al	O	$\Sigma\text{Cr,Cu,Fe,Mn,Ti}$
Ni-8Al ^a	91.69	8.30	0.010	~
Ni-15Al ^c	84.83	15.16	0.004	0.007
Ni-20Al ^a	80.17	19.81	0.009	0.004
Ni-23Al ^b	76.64	23.34	0.005	0.016
Ni-25Al ^b	74.61	25.38	0.004	0.003
Ni-27Al ^b	72.72	27.26	0.007	0.012
Ni-29Al ^a	70.23	29.76	0.008	0.002
Ni-30Al ^b	69.63	30.35	0.012	0.009
Ni-32Al ^a	67.6	32.4	0.006	~

fabricated by: a) NASA Glenn Research Center, b) Materials Preparation Center, Ames Laboratory, c) Air Force Research Laboratory-ML

Activity and Phase Equilibrium Measurements

The activities of Al and Ni in the various {alloy + Al_2O_3 } equilibria were determined by the vapor pressure method, where the partial pressure of characteristic vapor species, Al(g) and Ni(g) , in equilibrium with each sample, $p(i)$, are compared to a reference state, $p^\circ(i)$, according to: $a(i) = p(i) / p^\circ(i)$ [35-37]. The partial pressures were determined as a function of temperature by Knudsen effusion-cell mass spectrometry, *KEMS*, where $p(i)$ inside the effusion-cell is determined from the measured intensity of a representative ion beam, I_i , and absolute temperature, T , according to the relationship: $p(i) = I_i T / S_i$ (where S_i is the instrument sensitivity factor) [38]. The need to determine S_i and absolute partial pressure was removed in this study with a multiple effusion-cell vapor source (with 3 effusion-cells), *multi-cell KEMS*, which allows the relative partial pressure between samples in adjacent effusion-cells to be determined directly [39-42]. These measurements are made effectively independent of the effusion-cell by introducing two fixed apertures between the effusion-cell and ion-source (which define the molecular beam)

together with consistent alignment of each effusion-cell orifice with these fixed apertures. The multi-cell vapor source allows 2 alloys and pure-Au(s,l), as a secondary reference, to be measured in a single experiment with an indirect procedure, represented by Eq. 1. Activities are determined at each temperature, T , by comparing the ratio of the measured relative partial pressure of the characteristic vapor species ($p(\text{Al})$ and $p(\text{Ni})$) in equilibrium with the sample over $p^\circ(\text{Au})$ in equilibrium with pure-Au(s,l), $p(i) / p^\circ(\text{Au})$ or $I_i / I^\circ_{\text{Au}}$, to the tabulated ratio of $p^\circ(\text{Au})$ over the characteristic vapor species in equilibrium with the pure-element reference state $p^\circ(i)$, $[p^\circ(\text{Au}) / p^\circ(i)]$.

$$a(i) = \frac{p(i)}{p^\circ(\text{Au})} \cdot \left[\frac{p^\circ(\text{Au})}{p^\circ(i)} \right] = \frac{I_i}{I^\circ_{\text{Au}}} \cdot \frac{S_{\text{Au}}}{S_i} \cdot \frac{g_{\text{Au}}}{g_{\text{alloy}}} \cdot \left[\frac{p^\circ(\text{Au})}{p^\circ(i)} \right] \quad (\text{Eq. 1})$$

The $g_{\text{Au}} / g_{\text{alloy}}$ term accounts for variations in the effusion-orifices for a pair of cells [43-45] and was routinely measured to be 1.00 ± 0.01 in this study. The S_{Au} / S_i terms can be considered as instrument specific ionization cross-section ratios that relate the secondary reference, pure-Au(s,l) or $p^\circ(\text{Au})$, to the pure-element references, which are $\{\text{Al(l)} + \text{Al}_2\text{O}_3(\text{s})\}$ or $p^\circ(\text{Al})$ and $\{\text{Ni(s,l)} + \text{Al}_2\text{O}_3(\text{s})\}$ or $p^\circ(\text{Ni})$, respectively. These calibration factors, $S_{\text{Au}} / S_{\text{Al}}$ and $S_{\text{Au}} / S_{\text{Ni}}$, were measured in separate experiments as 0.154 ± 0.005 and 0.85 ± 0.03 , respectively and were independent of temperature. The tabulated ratio $[p^\circ(\text{Au}) / p^\circ(i)]$ in Eq. 1 was determined at each T by the 3rd-law treatment suggested by Paule [46] using the Gibbs free energy functions for the pure substances from reference [47] together with the measured 2nd-law reactions enthalpies listed in column 2 of Table 2. Because of the different ionization efficiency curves for Au(g), Al(g) and Ni(g) the same electron energy (25eV) was used for both S_{Au} / S_i and the alloy measurements.

Table 2. Reaction Enthalpies (at 298K) involving: Au(g), Ni(g), Al(g) and $\text{Al}_2\text{O(g)}$

Reaction	Measured (kJmol ⁻¹)	TPIS [47] (kJmol ⁻¹)
Au(s,l) = Au(g)	363.5±2.8	367.0±0.9 ^[46]
Ni(s) = Ni(g)	367.0±1.3*	428.0±8.0
Al(s) = Al(g)	428.3±2.6	330.0±3.0
4/3Al(s) + 1/3Al ₂ O ₃ (s) = Al ₂ O(g)	341.0±2.2	409.9±5.5
4/3Al(g) + 1/3Al ₂ O ₃ (s) = Al ₂ O(g)	414.2±3.6	-30.0±4.3
2Al(g) + O(g) = Al ₂ O(g)	-41.1±3.2	-1075.8±20.0
	-1075.5±9.0	

* 3rd law measurement

The *multi-cell KEMS* configuration also allows relative activities between two compositions, A relative to B , to be determined directly by measuring the relative partial pressure in equilibrium with the different samples in adjacent effusion-cells according to: $a(i)_{A-B} = p(i)_A / p(i)_B = I_i^A g^B / I_i^B g^A$. Again, the effect of variations in the effusion orifices were found to be minimal (*i.e.*, $g^B / g^A = 1.00 \pm 0.01$). Relative activities are a direct measure of the change in solution behavior and phase equilibrium between two compositions and remove any error from the thermodynamic properties of the pure-element reference states. In this study relative activity measurements focused on compositions Ni-23Al and Ni-27Al, either side of γ' -Ni₃Al in the $\{\gamma' + \gamma + \text{Al}_2\text{O}_3\}$ and $\{\gamma' + \beta + \text{Al}_2\text{O}_3\}$ three phase fields. In addition, a change in phase stability (phase boundary) is indicated by a change in the slope in a plot of the logarithm of $a(i)$ versus $1/T$, for both Al and Ni. The sensitivity of this measurement is improved by observing the ratio of the ion-intensities, $I_{\text{Ni}} / I_{\text{Al}}$, in the molecular beam from a single

effusion-cell with temperature^[49,50]. This technique was used to determine the phase boundaries and temperature of invariant reactions in this study.

The temperature was measured with a pyrometer (*Mikron Infrared, New Jersey, M190V-TS*) sighting a blackbody source (2.5 mm in diameter and 13.5 mm long) machined into the bottom of each effusion-cell and Mo-cell holder. Pure-Au(s,l) provides a primary temperature standard. The melting temperature of Au, $T_{mp} = 1337.33\text{K}$, which was used to calibrate the pyrometer in each experiment and the enthalpy of the sublimation of pure-Au(s,l) provided a systematic method of checking the accuracy of these results. The steady state or equilibrium condition in each effusion-cell was verified at each temperature with repeat measurements 30–45 minutes apart. The typical variation in temperature and ion-intensity between repeat measurements was less than 0.5 K and 1 %, respectively. Measurements were made at a range of temperatures over three days and were taken in a random order.

Results

This publication reports the results from thirteen separate experiments. In each experiment the enthalpy of sublimation of pure-Au(s,l), $\Delta H_{sub}\text{Au}(298)$, was determined and used as a check for consistent measurement procedure. In all cases the 2nd and 3rd law values were within the range $362.8 \pm 3.0 \text{ kJmol}^{-1}$ and $367.8 \pm 1.0 \text{ kJmol}^{-1}$, respectively, and agree with the values listed in Table 2. The activities of Al and Ni of all nine alloys are plotted as the logarithm of activity versus $1/T$ and are shown in Figs. 1 and 2, respectively. For the composition range 8 – 32 at.% Al, the Al activity varied by about a factor of 300 while the Ni activity varied by about a factor of 2. The relative errors in both sets of data are similar, however, the different scales make the $a(\text{Al})$ data look more precise than $a(\text{Ni})$ data. The sensitivity of

these measurements to changes in sample composition relates directly to the Gibbs free-energy surface. In this case a difference of 1 at.% Al is observable. In Figs. 1 and 2 the univariant phase fields, {phase^I + phase^{II} + Al_2O_3 }, in the Ni-Al-O system exist as lines and are identified by solid-lines labeled with stable alloy phases $\gamma' + \gamma$, $\gamma' + \beta$, $\gamma + \beta$, $L + \gamma$ and $L + \beta$. In regions where bivariant phase fields {phase + Al_2O_3 } exist the behavior of each sample is identified by a dashed line and the stable alloy phase is identified. In the univariant phase fields the composition of the condensed phases are fixed but unknown, while in the bivariant fields the composition of the alloy phase is known (oxygen content is unknown) and listed in Table 1.

In Fig. 1 and 2 the changes in slope observed in a particular plot correspond to the sample crossing a phase boundary. In this way these results provide important information about the equilibrium phase boundaries and invariant transformations in the Ni-Al-O system. The most striking feature of these results is the four univariant phase fields ($\gamma' + \gamma$, $\gamma' + \beta$, $L + \gamma$ and $L + \beta$) converge within the temperature range $T = 1630 \sim 1640 \text{ K}$, where the same Al and Ni activities were measured for Ni-23Al, Ni-25Al and Ni-27Al samples (note: $a(\text{Al}) \neq a(\text{Ni})$). The convergence of the $\gamma' + \gamma$ and $\gamma' + \beta$ and the subsequent divergence of the $L + \gamma$ and $L + \beta$ univariant phase fields is due to changes in the composition of γ -(Ni), γ' -Ni₃Al, β -NiAl and L with T , but measuring the same Al and Ni activities indicate Ni-23Al, Ni-25Al and Ni-27Al enter the same phase field, which is inconsistent with the current Ni-Al phase diagram^[21-25]. If the current Ni-Al phase diagram is correct γ' -Ni₃Al remains stable up to the liquidus, and compositions either side of γ' -Ni₃Al (*i.e.*, Ni-23Al and Ni-27Al) must remain in different phase fields, separated by the composition range of γ' -Ni₃Al below the liquidus and by L above the liquidus. These new results clearly indicate that this is not the case and the Ni-Al phase diagram needs to be revised.

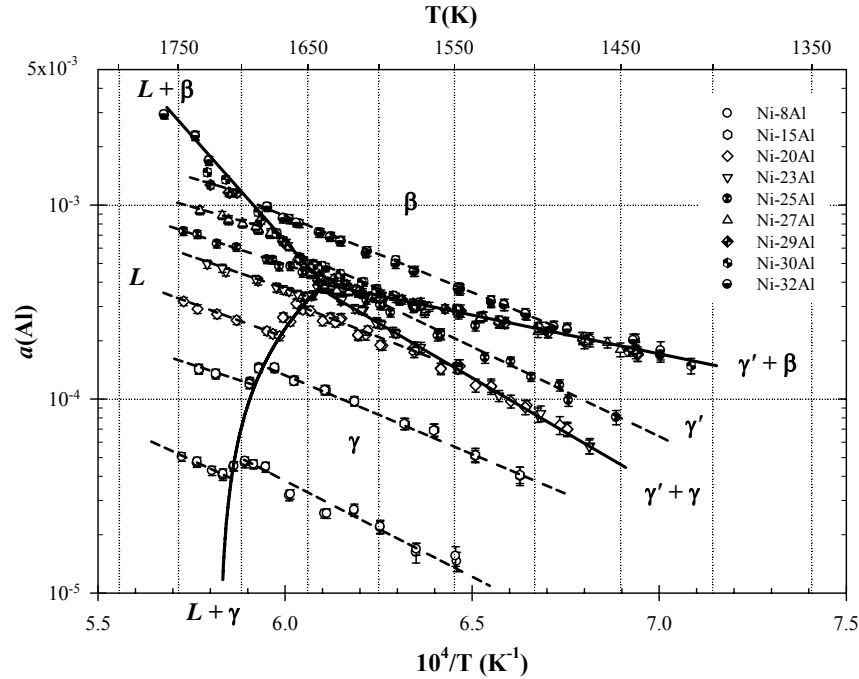


Figure 1. A plot of $\log a(\text{Al})$ vs. $1/T$ for alloys: Ni-8Al, Ni-15Al, Ni-20Al, Ni-23Al, Ni-25Al, Ni-27Al, Ni-29Al, Ni-30Al, Ni-32Al.

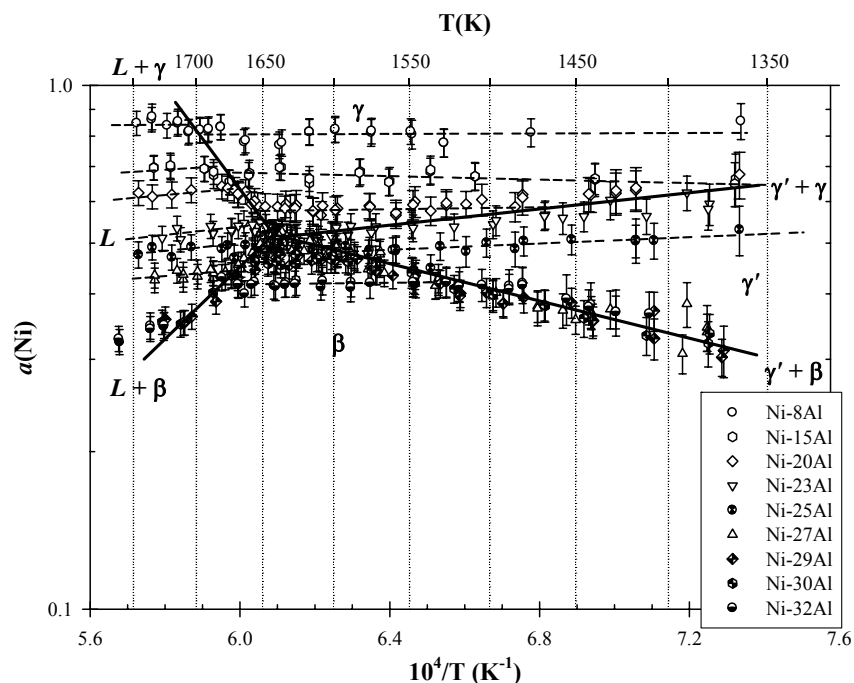


Figure 2. A plot of $\log a(\text{Ni})$ vs. $1/T$ for alloys: Ni-8Al, Ni-15Al, Ni-20Al, Ni-23Al, Ni-25Al, Ni-27Al, Ni-29Al, Ni-30Al, Ni-32Al.

More detailed information about the phase equilibria involving γ' - Ni_3Al was obtained by placing Ni-27Al and Ni-23Al samples in adjacent effusion-cells and measuring the relative activities of Ni and Al (for Ni-27Al / Ni-23Al). Repeated measurements were combined and the data are plotted in Fig. 3, as relative activity versus T . This plot shows the relative activity of Al and Ni

converge to unity over the temperature range $T = 1633 \pm 1 - 1640 \pm 1$ K where both alloys must be in the same phase field. Below 1633 K alloy Ni-27Al and Ni-23Al are in the $\gamma' + \beta$ and $\gamma' + \gamma$ univariant phase fields and the relative activities effectively represent the chemical potential difference across the composition range of γ' - Ni_3Al , which decreases with increasing T .

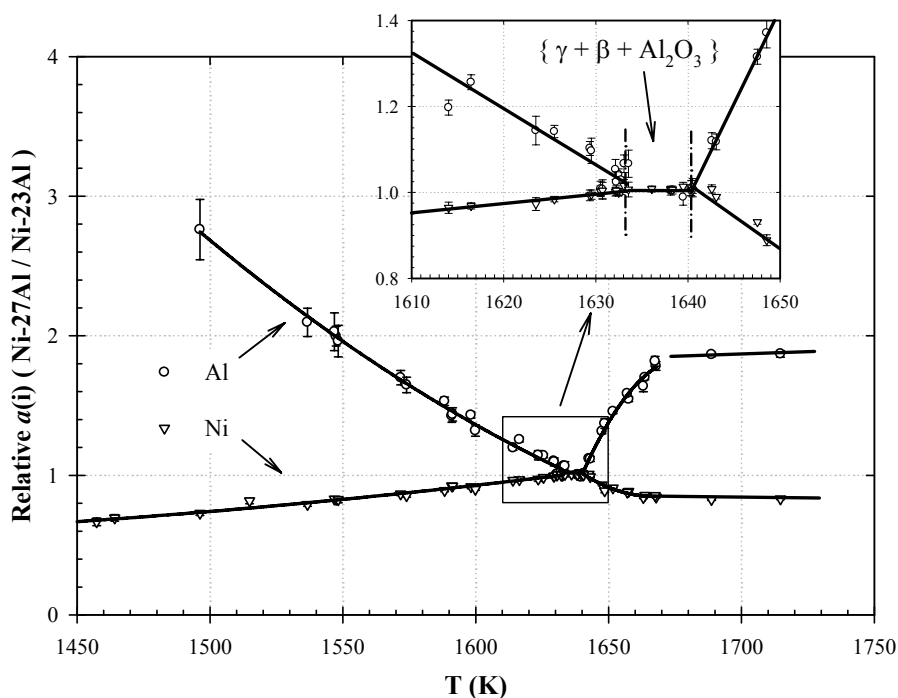


Figure 3. Relative $a(\text{Al})$ and $a(\text{Ni})$ for Ni-27Al / Ni-23Al vs. T (K).

At $T = 1633$ K this difference goes to zero (*i.e.*, relative $a(\text{Al})$ and relative $a(\text{Ni}) = 1$) signifying that γ' -Ni₃Al is no longer stable for $T > 1633$ K. Similarly above $T = 1640$ K, but below the liquidus, the alloys are in the $L + \beta$ and $L + \gamma$ univariant phase fields and the relative activities represent the difference across the L phase. For $T < 1640$ K the difference also goes to zero and indicating L is no longer stable. According to the Gibbs phase-rule ($f = c + 2 - p$), a maximum of 5 phases can coexist at an invariant point in the Ni-Al-O system. Therefore only 4 phases can be stable over the temperature range $T = 1633 - 1640$ K and there must be an invariant point at either end of this temperature range. As γ -(Ni) and β -NiAl both remain stable up to the liquidus, γ' -Ni₃Al must decompose to $\gamma + \beta$ on heating via the peritectoid reaction, $\gamma + \beta (+ \text{Al}_2\text{O}_3) = \gamma' (+ \text{Al}_2\text{O}_3)$, at $T = 1633 \pm 1$ K, which corresponds to the $\{\gamma' + \gamma + \beta + \text{Al}_2\text{O}_3\}$ equilibrium. Similarly at $T = 1640 \pm 1$ K the liquid decomposes on cooling via the eutectic reaction, $L (+ \text{Al}_2\text{O}_3) = \gamma + \beta (+ \text{Al}_2\text{O}_3)$, which corresponds to the $\{L + \gamma + \beta + \text{Al}_2\text{O}_3\}$ equilibrium. Based on this argument the $\{\gamma + \beta + \text{Al}_2\text{O}_3\}$ equilibrium is stable over the range $T = 1633 - 1640$ K.

Information about the broader phase equilibria in the Ni-Al-O system was obtained from the measured ratio of the ion-intensities, $I_{\text{Ni}} / I_{\text{Al}}$, in the molecular beam from each effusion-cell with T . This is effectively the ratio, $a(\text{Ni}) / a(\text{Al})$, in each alloy and is independent of results shown in Figs. 1, 2 and 3. Figure 4 shows the results from all samples on a plot of the natural logarithm of $I_{\text{Ni}} / I_{\text{Al}}$ versus $1/T$. The same procedure, discussed above, was used to identify the behavior of each sample as it traversed the univariant and bivariant phase fields. This data is more suited to identifying phase boundaries and invariant transformations. For example, point A in Fig. 4 represents the solidus $\gamma + \text{Al}_2\text{O}_3 / L + \gamma + \text{Al}_2\text{O}_3$ ($T = 1675 \pm 5$ K) for the Ni-15Al sample, while point B represents the γ -(Ni) solvus $\gamma + \gamma' + \text{Al}_2\text{O}_3 / \gamma + \text{Al}_2\text{O}_3$ ($T = 1541 \pm 8$ K) for Ni-20Al. An invariant

transformation is identified by the point of intersection of 2 univariant phase fields. All the phase boundary data obtained in this study (from Fig. 3 and 4) are summarized in Table 3.

Table 3. Phase Boundaries and Invariant Reactions

Alloy	Phase Boundary	T (K)
Ni-8Al	$\gamma + \text{Al}_2\text{O}_3 / L + \gamma + \text{Al}_2\text{O}_3$	1697 ± 4
	$L + \gamma + \text{Al}_2\text{O}_3 / L + \text{Al}_2\text{O}_3$	1715 ± 3
Ni-15Al	$\gamma + \text{Al}_2\text{O}_3 / L + \gamma + \text{Al}_2\text{O}_3$	1675 ± 5
	$L + \gamma + \text{Al}_2\text{O}_3 / L + \text{Al}_2\text{O}_3$	1695 ± 5
Ni-20Al	$\gamma + \gamma' + \text{Al}_2\text{O}_3 / \gamma + \text{Al}_2\text{O}_3$	1541 ± 8
	$\gamma + \text{Al}_2\text{O}_3 / L + \gamma + \text{Al}_2\text{O}_3$	1651 ± 4
Ni-23Al	$L + \gamma + \text{Al}_2\text{O}_3 / L + \text{Al}_2\text{O}_3$	1674 ± 3
	$\gamma + \gamma' + \text{Al}_2\text{O}_3 / \gamma + \beta + \text{Al}_2\text{O}_3$	1633 ± 1
Ni-25Al	$\gamma + \beta + \text{Al}_2\text{O}_3 / L + \gamma + \text{Al}_2\text{O}_3$	1640 ± 1
	$L + \gamma + \text{Al}_2\text{O}_3 / L + \text{Al}_2\text{O}_3$	1649 ± 1
Ni-27Al	$\gamma' + \text{Al}_2\text{O}_3 / \gamma' + \beta + \text{Al}_2\text{O}_3$	1614 ± 6
	$\gamma' + \beta + \text{Al}_2\text{O}_3 / \gamma + \beta + \text{Al}_2\text{O}_3^\dagger$	1633 ± 1
Ni-29Al	$\gamma + \beta + \text{Al}_2\text{O}_3 / L + \beta + \text{Al}_2\text{O}_3$	1639 ± 2
	$L + \beta + \text{Al}_2\text{O}_3 / L + \text{Al}_2\text{O}_3$	1646 ± 1
Ni-30Al	$\gamma' + \beta + \text{Al}_2\text{O}_3 / \gamma + \beta + \text{Al}_2\text{O}_3$	1633 ± 1
	$\gamma + \beta + \text{Al}_2\text{O}_3 / L + \gamma + \text{Al}_2\text{O}_3$	1640 ± 1
Ni-32Al	$L + \beta + \text{Al}_2\text{O}_3 / L + \text{Al}_2\text{O}_3$	1672 ± 1
	$\gamma' + \beta + \text{Al}_2\text{O}_3 / \beta + \text{Al}_2\text{O}_3$	1605 ± 5
Ni-32Al	$\beta + \text{Al}_2\text{O}_3 / L + \beta + \text{Al}_2\text{O}_3$	1642 ± 3
	$L + \beta + \text{Al}_2\text{O}_3 / L + \text{Al}_2\text{O}_3$	1704 ± 2
Ni-30Al	$\gamma' + \beta + \text{Al}_2\text{O}_3 / \beta + \text{Al}_2\text{O}_3$	1585 ± 8
	$\beta + \text{Al}_2\text{O}_3 / L + \beta + \text{Al}_2\text{O}_3$	1657 ± 2
Ni-32Al	$L + \beta + \text{Al}_2\text{O}_3 / L + \text{Al}_2\text{O}_3$	1714 ± 5
	$\gamma' + \beta + \text{Al}_2\text{O}_3 / \beta + \text{Al}_2\text{O}_3$	1485 ± 4
Ni-32Al	$\beta + \text{Al}_2\text{O}_3 / L + \beta + \text{Al}_2\text{O}_3$	1681 ± 4

† It was impossible to determine the $\gamma' + \beta + \text{Al}_2\text{O}_3 / \gamma + \beta + \text{Al}_2\text{O}_3$ transformation temperature from these data.

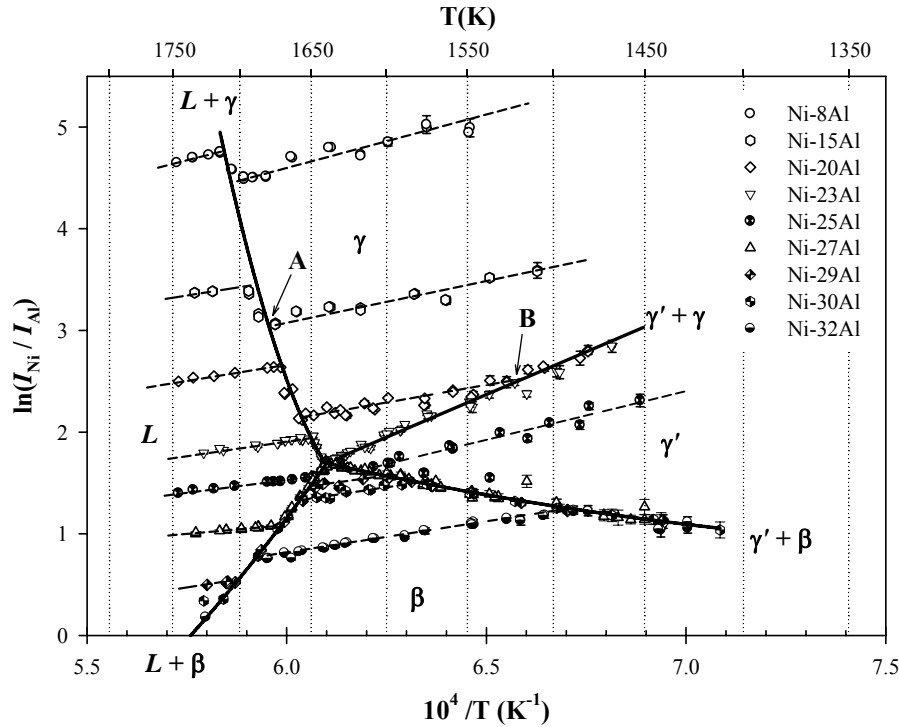


Figure 4. Measured $\ln(I_{\text{Ni}} / I_{\text{Al}})$ vs. $1/T$ for: Ni-8Al, Ni-15Al, Ni-20Al, Ni-23Al, Ni-25Al, Ni-27Al, Ni-29Al, Ni-30Al, Ni-32Al.

Discussion

The measured reaction enthalpies for $\{\text{Al(l)} + \text{Al}_2\text{O}_3\}$ listed in column 2 of table 2 are significant with potentially wide ranging ramifications. While the vaporization behavior of $\{\text{Al(l)} + \text{Al}_2\text{O}_3\}$ has no influence on the eutectic and peritectoid transformations involving γ' -Ni₃Al observed in this study, it raises some questions about the accepted thermodynamic properties of Al(s,l,g), the value of $p^\circ(\text{Al})$ and thus previously measured $a(\text{Al})$ data and published thermodynamic description of systems that contain Al. Therefore a brief discussion of these new results and review of the previously published thermodynamic data for Al is provided. The reaction enthalpies listed in table 2 were determined during repeated measurements of the $S_{\text{Au}} / S_{\text{Al}}$ and $S_{\text{Au}} / S_{\text{Al}_2\text{O}_3}$ calibration terms that relate Au(s,l) or $p^\circ(\text{Au})$, to $\{\text{Al(l)} + \text{Al}_2\text{O}_3(\text{s})\}$ or $p^\circ(\text{Al})$ and $p^\circ(\text{Al}_2\text{O}_3)$. In these experiments the 2nd and 3rd law values for the enthalpy of sublimation of Au(s) were measured in the ranges $363.0 \pm 3.0 \text{ kJmol}^{-1}$ and $367.0 \pm 1.5 \text{ kJmol}^{-1}$, respectively. Similarly the 2nd law enthalpy of sublimation of Al(s) was repeatedly measured in the range $341.0 \pm 2.2 \text{ kJmol}^{-1}$. (The Gibbs free energy functions ($G^\circ(T) - H^\circ(298)/T$) for the pure substances used in these calculations came from references 46 and 47.) This differs significantly from the currently recommended value of $330.0 \pm 3.0 \text{ kJmol}^{-1}$ [47,48] but given that the accepted sublimation of Au(s) was reproduced in the same experiments there is no compelling reason to question the validity of these new results. While some issues were investigated by the author [51,52] the underlying fact is that Al(l) is highly reactive and to date the most suitable container has not been widely used, and to make matters worse different containers have been used for vaporization and calorimetric studies. A review of the literature reveals that graphite, BeO, TaC, Ta₃Si, ZrC, BN, TiB₂, Al₂O₃ and ZrO₂ have all been used to contain Al(l), with varying degrees of success, in melting point, calorimetry and vaporization studies. Because Al₂O₃ is such a stable compound it is impractical to consider Al independently of O, however, the Al-O phase diagram [33] indicates Al₂O₃ is in equilibrium with Al(l) and the O saturation limit is less than 0.1at% O at temperatures above 2000K. In addition, Al₂O₃ exhibits slow diffusion kinetics for both O and Al. Accordingly Al₂O₃ is the perfect container for Al(l) and the univariant equilibrium $\{\text{Al(l)} + \text{Al}_2\text{O}_3\}$ unambiguously defines $p^\circ(\text{Al})$, $p^\circ(\text{Al}_2\text{O}_3)$ and $p^\circ(\text{O})$ and provides an excellent experimental reference state for Al, Al₂O₃, O and Al₂O₃ [43].

Unfortunately to date there have been no reliable studies reporting the vaporization behavior of $\{\text{Al(l)} + \text{Al}_2\text{O}_3\}$. As a result all the studies cited in the thermodynamic compilations (JANAF, TPIS, CODATA, etc.) as the basis of the recommended $\Delta H_{\text{sub}}\text{Al}(298)$ value used BeO(s) as a container for Al(l) [53-55]. This is surprising because both Brewer *et al.* [54] and Potter *et al.* [53] clearly state that Al(l) reacts with the BeO(s) to form unidentified AlBe_xO_y compounds together with the dissolution of 2 ~ 5at% Be in Al(l). If these observations are correct the reported vaporization behavior cannot be assigned to pure Al(l) and indeed this is the likely cause for the discrepancy seen in table 2. Further, reference [53] is a progress report for a wide ranging project that only contains a brief summary of the vaporization of Al(l) consisting of the 2nd and 3rd law values for $\Delta H_{\text{sub}}\text{Al}(298)$ and it turns out that the measured $p(\text{Al})$ versus T data was never actually published [56]. The critical calorimetric study of Al(l) by McDonald [57] used to determine the heat capacity and Gibbs free energy function of Al(l) used both BN and TiB₂ to contain Al(l). BN is unsuitable because AlN and AlB₂ need to form before Al(l)

is stable but neither compound provides a suitable diffusion barrier to stop the reaction of Al with the BN container. After Al₂O₃, TiB₂ is probably the best container for Al(l) because the $\{\text{Al(l)} + \text{TiB}_2\}$ and $\{\text{Al(l)} + \text{TiAl}_3 + \text{TiB}_2\}$ equilibria exist, in fact TiB₂ has a long history as a grain refiner for Al-based alloys. However, at high T significant dissolution of Ti and B in Al(l) occurs and it is uncertain if a univariant equilibrium exists. It is unclear why calorimetric / heat capacity measurements of Al(l) were never made in Al₂O₃ containers. Given that Al is such an important component in a wide range of alloys currently used in many industrial applications there is an urgent need to re-measure the heat capacity and vaporization behavior of Al(l) to elucidate the inconsistency in the accepted thermodynamic data. Ideally heat capacity and vaporization measurements of Al(l) need to be made in consistent containers (*i.e.*, Al₂O₃, TiB₂ and possibly AlN).

On returning to the eutectic and peritectoid reactions observed in this study, it is important to recognize the system considered in this study is essentially the same as in all previous studies [1-25]. While the system was identified as Ni-Al-O in this study, changing the description of the system cannot influence the actual measurements. The different behavior reported here is a consequence of the isothermal conditions, obtaining equilibrium at each T and a significant improvement in experimental capability provided by the *multi-cell* KEMS technique. The limited temperature range where $\{\gamma + \beta + \text{Al}_2\text{O}_3\}$ is stable, the similar composition of the L in the eutectic and γ' -Ni₃Al (as shown in Fig. 5) together with fast reaction kinetics make the details of this behavior very difficult to observe with more conventional non-isothermal techniques. The ability to compare the partial thermodynamic properties from different samples at temperature provides direct information about differences in the phase equilibria and provides an excellent tool for investigating complex phase transformations under equilibrium conditions. This capability needs to be applied to a wide range of technically important alloy and ceramic systems.

Combining the data listed in tables 1 and 3 provide 24 (X_{Al} , T) points that correspond to various phase boundaries, which were used to construct the section of the "Ni-Al" phase diagram shown in Fig. 5. Given that measurements made in the Ni-Al-O system have been routinely applied to the Ni-Al system, for the purpose of this discussion the current results are also presented in this way to allow a direct comparison with previous studies. To avoid confusion, however, Fig. 5 is a projection of the alloy phase boundaries observed in the Ni-Al-O system on to the Ni-Al binary edge of the ternary system, where the alloy phases are saturated with O and in equilibrium with Al₂O₃. For the purpose of this comparison the proposed phase equilibrium (the thick solid / dashed lines and data points) has been superimposed on the current Ni-Al phase diagram (thinner solid grey lines) from the assessment of Huang and Chang [24]. There is just enough data to define the shape of the liquidus, solidus (γ -(Ni), β -NiAl and γ -(Ni) + β -NiAl) and β -NiAl solvus boundaries shown in Fig. 5, but the placement of the γ -(Ni) solvus and γ' -Ni₃Al phase boundaries is pure conjecture based on the currently accepted behavior. The critical finding of this study is that γ' -Ni₃Al forms via the peritectoid reaction $\gamma + \beta (+ \text{Al}_2\text{O}_3) = \gamma' (+ \text{Al}_2\text{O}_3)$, at $T = 1633 \pm 1 \text{ K}$, below the liquidus and γ' -Ni₃Al is not involved in either a eutectic or peritectic reaction as currently accepted.

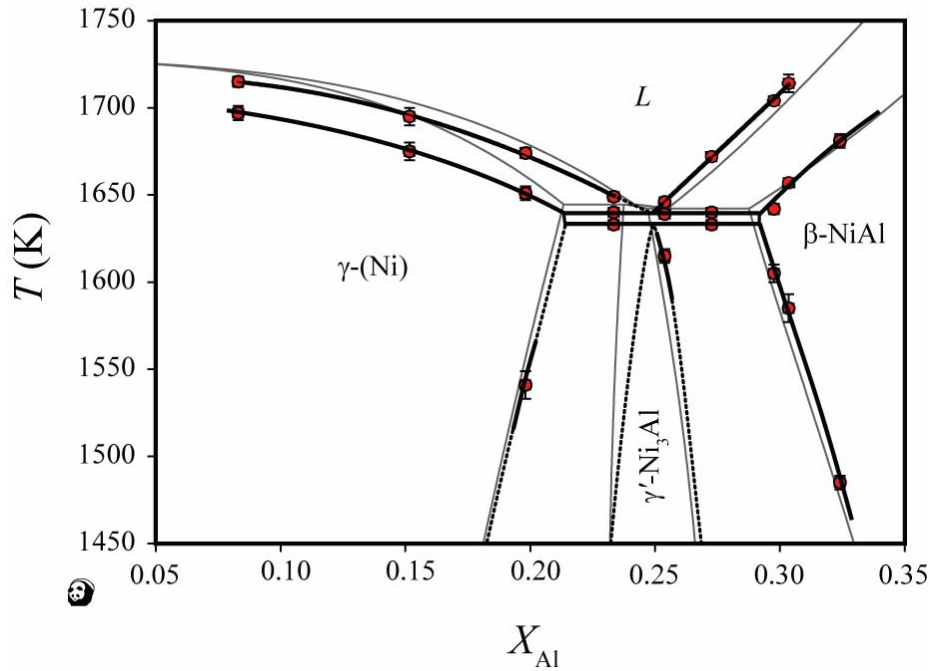


Figure 5. The projection of the alloy phase boundaries within the Ni-Al-O system, in equilibrium with Al_2O_3 and saturated with O, on to the Ni-Al binary edge. The phase boundary points in this figure were taken from the data listed in Tables 1 and 3. These new data are superimposed (shown as thick solid / dashed lines) over the current Ni-Al phase diagram (thinner solid grey line) from the assessment of Huang and Chang ^[24].

Above the peritectoid, over the temperature range $T = 1633 - 1640$ K, γ -(Ni) and β -NiAl are in equilibrium and at $T = 1640 \pm 1$ K there is a eutectic reaction, $L (+ \text{Al}_2\text{O}_3) = \gamma + \beta (+ \text{Al}_2\text{O}_3)$. An estimate of the liquid composition at the eutectic, 24.8 ± 0.2 at% Al, was obtained by extrapolating the measured liquidus on either side of the eutectic down to the eutectic temperature, $T = 1640$ K. The composition of γ' -Ni₃Al at the peritectoid reaction cannot be determined from these data but it is clearly to the Ni-rich side of Ni-25Al composition and close to the composition of L in the eutectic. Apart from the solidification behavior of γ' -Ni₃Al there is good agreement in the γ -(Ni) solvus, β -NiAl solvus, liquidus, β -NiAl solidus and Al-rich γ' -Ni₃Al phase boundary between the results reported in this study and Huang's assessment. However, there is a significant difference in the position of the γ -(Ni) solidus and nature of this discrepancy needs further investigation.

In view of the fact that these results represent such a large departure from the accepted phase equilibrium in the Ni-Al system, the reproducibility of this behavior was checked with three independent measurements (absolute activities Fig. 1 and 2, relative activities Fig. 3 and ion-intensity ratios Fig. 4). All three sets of measurements showed consistent behavior across the thirteen experiments. It would be ideal if the $\{\gamma + \beta + \text{Al}_2\text{O}_3\}$ phase field was confirmed directly by quenching in the high temperature microstructure and observing it by optical and electron microscopy. Unfortunately this experiment cannot be conducted with the *multi-cell* furnace because it is under a high vacuum and high cooling rates are impossible. A cursory review of the relevant literature reveals that $\gamma + \beta$ eutectic structures have

in fact been routinely observed in the Ni-Al system and the factors affecting their formation, during directional solidification, have been studied in detail ^[12,14,16]. These studies were founded on view that the current Ni-Al phase diagrams is correct (*i.e.*, γ' -Ni₃Al is stable up to the liquidus and is involved in both $L + \gamma = \gamma'$ and $L = \beta + \gamma'$ reactions) and the aim was to understand how the "metastable" $\gamma + \beta$ eutectic structure formed as a function of directional solidification conditions. However, in the light provided by these new results and the proposed phase equilibrium in the Ni-Al system it is interesting to reconsider the discussion surrounding the stability and formation of $\gamma + \beta$ eutectic structures.

Lee and Verhoeven ^[12-15] used the Bridgman technique with a thermal gradient of 20 – 40 K/cm in the liquid. The alloy samples (with Al content 23 – 27 at% Al) were placed in an Al_2O_3 tube and a sample feed rate of 0.8 – 61 $\mu\text{m/s}$ was used. A range of quenching methods (in order of increasing cooling rate: air, water, ice-water and molten Ga-In) were employed in an attempt to freeze in the steady-state solidification structures. In addition to the $\gamma + \beta$ eutectic these studies also revealed unusually wide lamella spacing in $\gamma' + \beta$ structures, γ' / γ' interfaces ^[13] and two cellular structures of $\gamma' + \beta$ and $\gamma + \gamma' + \beta$ (the latter contradicts the Gibbs phase rule), all of which the authors felt could not be adequately explained. While their initial publication ^[12] is widely cited as a critical study proving the accepted Ni-Al phase equilibrium in the development of the thermodynamic descriptions ^[22-24] all subsequent publications reporting unusual microstructures were and are still inexplicably ignored. From this

author's perspective, the basic finding was directional solidification conditions did not influence which eutectic formed. Rather, the quenching method determined if a $\gamma + \beta$ or $\gamma' + \beta$ was observed. This conclusion comes from the idea that reactions occurring at temperature during the steady state directional solidification are independent of any subsequent reactions occurring during the quenching and the probability of quenching-in the steady-state solidification structure increases as the cooling rate increases. The $\gamma + \beta$ eutectic structure was observed in samples that were cooled at the highest rates (*i.e.*, a molten Ga-In bath and sometimes iced-water quench) while the various $\gamma' + \beta$ and $\gamma' + \gamma$ structures were only observed in samples cooled more slowly. This indicates the $\gamma + \beta$ eutectic structure must be the stable steady-state solidification structure and not a metastable structure as has been suggested. The formation of a $\gamma + \beta$ eutectic under steady-state solidification conditions agrees with the new Ni-Al phase diagram proposed in this study. The most compelling evidence of this is found in figure 2 of reference [14] which presents a quenched high-temperature microstructure that clearly shows: the $L = \gamma + \beta$ eutectic at the liquid / solid interface; the $\gamma + \beta$ eutectic structure over a limited temperature range; the $\gamma + \beta = \gamma'$ peritectoid with γ' nucleating at γ / β interfaces and growth into lamella of γ and β as the sample continues to cool (producing metastable $\gamma + \gamma' + \beta$ structures).

In a similar study, Hunziker and Kurz^[16] investigated the factors influencing the formation of $\gamma + \beta$ eutectic structures over the composition range 11 – 30 at.%Al. They used a laser surface resolidification technique on samples at room temperature and 1273 K over a wide range of velocities (5×10^{-4} – 0.5 ms^{-1}). In all cases the un-melted bulk of these samples provide a large enough heat-sink to give a high cooling rate (or auto-quenching) which successfully preserved the steady-state solidification structure. Only the $\gamma + \beta$ eutectic alone and together with prior- β or prior- γ were observed. The absence of primary γ' -Ni₃Al indicates it is not stable up to the liquidus and it was suggested that γ' form by a nucleation and growth process (shown schematically in figure 9 in reference [16]) that is effectively a $\gamma + \beta = \gamma'$ peritectoid reaction. This behavior agrees with Lee and Verhoeven and the results presented in this publication. Further independent confirmation of this phase equilibrium is also found in the measurements of the temperature dependence of the long-range order parameter of γ' -Ni₃Al made by Bremer *et al.* [7]. Their study used high temperature γ -ray diffraction to measure the Bragg scattering intensity of the fundamental (200) and superlattice (100) reflection to calculate the long range order in Ni-24Al and Ni-22Al alloys. The results clearly show γ' -Ni₃Al remains highly ordered up to about $T = 1590 \text{ K}$ but starts to disorder above $T = 1610 \text{ K}$ and is completely disordered below the solidus. Bremer *et al.* suggest an order-disorder transition but these results could also indicate the $\gamma + \beta = \gamma'$ peritectoid reaction identified in the current study.

If all the reported invariant transformation behavior is viewed objectively with the possibility of choosing between the current Ni-Al phase diagram and the behavior shown in Fig. 5 it is clear that all the previously reported data are completely consistent with the new behavior reported in this study. However, if the phase equilibrium reported in this study is used all of these prior observations can be explained in a straightforward manner.

Conclusion

The activities of Al and Ni were measured with the *multi-cell* KEMS technique over the composition range 8 – 32 at.%Al and temperature range and $T = 1400 - 1750 \text{ K}$ to better understand the equilibrium solidification behavior of γ' -Ni₃Al containing alloys in the Ni-Al-O system. The data clearly showed the univariant phase fields: $\{\gamma' + \gamma + \text{Al}_2\text{O}_3\}$, $\{\gamma' + \beta + \text{Al}_2\text{O}_3\}$, $\{\gamma + L + \text{Al}_2\text{O}_3\}$ and $\{\beta + L + \text{Al}_2\text{O}_3\}$, converged over the temperature range $T = 1633 \pm 1 - 1640 \pm 1 \text{ K}$ where the same $a(\text{Al})$ and $a(\text{Ni})$ exist in the composition range 23 – 27 at.%Al. This behavior reveals that γ' -Ni₃Al is not involved in either a eutectic or peritectic but forms via the peritectoid reaction $\gamma + \beta (+ \text{Al}_2\text{O}_3) = \gamma' (+ \text{Al}_2\text{O}_3)$, at $T = 1633 \pm 1 \text{ K}$. The $\{\gamma + \beta + \text{Al}_2\text{O}_3\}$ phase field is stable over the temperature range $T = 1633 - 1640 \text{ K}$ and equilibrium solidification occurs by the eutectic reaction, $L (+ \text{Al}_2\text{O}_3) = \gamma + \beta (+ \text{Al}_2\text{O}_3)$ at $T = 1640 \pm 1 \text{ K}$ with a liquid composition of $24.8 \pm 0.2 \text{ at.\%Al}$ (at an unknown O content). This behavior is inconsistent with the current Ni-Al phase diagram in the vicinity of γ' -Ni₃Al and a new Ni-Al phase diagram (projected from the Ni-Al-O system) was proposed based on these new results. This new Ni-Al phase diagram provides a compelling explanation for some steady-state solidification structures that are routinely observed but are difficult to explain in terms of the current Ni-Al phase diagram. If viewed objectively, this new phase diagram is consistent with all previously reported observations and provides a much simpler reaction scheme.

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